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12/19/2003. Author: Ted Gibb, Lucille
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Solution: The concentration of unreacted ethanoic acid, $\text{CH}_3\text{COOH}(\text{aq})$, can be determined:

$$\text{CH}_3\text{COOH}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{CH}_3\text{COO}^-\text{Na}^+(\text{aq}) + \text{H}_2\text{O}(\text{l})$$

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Solution: OH OH OH 0.50 mmol mL --
-- = x = n c V x 25.0 mL OH n - = 12.5
mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5
mmol = = n V 0.10 mmol HCl(aq) /mL V
= 120 mL 2. (a) Given: V HCl(aq) = 25.00
mL; [HCl(aq)] = 0.350 mol/L; V

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NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H+(aq)], pH Analysis: Before titration, amount of H+ is: H(aq) H(aq) 0.350 mmol mL ++ +=x = nc V

Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547

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