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$V(2) = m kT P(3)$

where m is the average
mass of an air
molecule, N is the

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number of molecules in volume V and T is the temperature. We therefore get a differential equation for the pressure: $dP/dz = -\rho g$ (4) $= -mg/kT P$ (5) This is called the barometric equation. In a realistic model, both g and

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