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the intermolecular attractions resulting from the constant motion of electrons and the creation of instantaneous dipoles; forces act between all atoms AND molecules; forces are NOT present in intramolecular forces; increases with atomic mass or molar mass.

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Organic Chemistry Chapter 6 and 7. Henry Lowe. 16 October 2020. 71 test answers. question. alkyl halide. answer. halogen atom bonded to an sp³ hybridized carbon atom that is tetrahedral. question.

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molar mass of C₆H₆O = 94.11 g 21. a. molar mass of SO₃ = 80.07 g 49.2 mg SO₃ 1 g 1000 mg 1 mol 80.07 g = 6.14 10⁻⁴ mol SO₃ b. molar mass of PbO₂ = 239.2 g 7.44 x 10⁴ kg PbO 2 1000 g 1 kg 1 mol 239.2 g = 3.11 10⁵ mol PbO 2 c. molar mass of CHCl₃ = 119.37 g 59.1 g CHCl₃ 1 mol 119.37 g = 0.495 mol CHCl₃ d. molar mass of C₂H₃Cl₃ = 133.39 g 3.27 ...

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Modern Chemistry Chapter 6 Chemical Bonding . Chemical Bond •A link between atoms that results from the mutual attraction of their nuclei for electrons -Electrostatic attraction between proton and electron -Classified by the way the valence e- are distributed around nuclei of combined atoms .

Modern Chemistry Chapter 6 - Harrison County Schools

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

6 Chemical Bonding - Effingham County Schools / Overview

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AP Chemistry Practice Test, Ch. 6: Thermochemistry MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. DH 4Al (s) + 3O₂ (g) → 2Al₂O₃ (s) A)exothermic, released B)exothermic, absorbed.

AP Chemistry Practice Test, Ch. 6: Thermochemistry ...

Chapter 6 - Thermochemistry Chapter 7 (Part 1) - Atomic Structure & Periodicity Chapter 7 (Part II) & Chp. 19 - Atomic Structure & Periodicity, Nuclear Chemistry

Chapter 6 - Thermochemistry - Mrs. Duffey - FHN

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