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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the

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balanced chemical equation ...

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Chapter 12 "Stoichiometry" Mr. Mole
Section 12.1 The Arithmetic of Equations

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OBJECTIVES: Explain how balanced equations apply to both chemistry and everyday life. Section 12.1 The Arithmetic of Equations **OBJECTIVES:** Interpret balanced chemical equations in terms of: a) moles, b) representative particles, c) mass, and d) gas volume (Liters ...

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Chapter 12 “Stoichiometry”

Chapter 12 - Stoichiometry. Homework.

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Drawings of Reactions.

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sections 12.2 making notes of new
things. Write at least 3 questions in your
Chemistry Journal; Ask your questions!
Worksheet: Nuts and Bolts:
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Defining Stoichiometry pages 368–372

Practice Problems pages 371–372 1.

Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is

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