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Section 16.1 - The Nature of Covalent Bonding Practice Problems 1. Draw electron dot structures for each molecule. a. chlorine b. bromine c. iodine What do you observe about the three structures? They all look similar with respect to the electrons. 2. The

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following molecules have single covalent bonds. Draw an electron dot structure for each. a. H_2O_2 b. PCl_3

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Ionic bonding occurs between a cation and an anion while a covalent bond occurs between two nonmetals. In an ionic bond, one electron is

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taken from another atom and is completely removed. In a covalent bond, the electron is shared between two molecules. An ionic bond forms an ionic compound while a covalent bond forms a molecule.

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In covalent bonds the electrons are not

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always shared equally
by the atoms Atom
with higher EN value
will pull shared
electrons closer to it
Covalent bonds can be:
Polar: unequal sharing
of electrons Nonpolar:
equal sharing of
electrons Look at EN
difference (Table 14.2
on p. 405) to
determine bond type
Bond type: Nonpolar
covalent: EN diff of ...

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Bonding Answer

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Bonding Chapter 16: 1

- 26; 28, 30, 31, 35-37,

40, 43-46, Extra Credit:

50-53, 55, 56, 58, 59,

62-67 Section 161 -The

Nature of Covalent

Bonding Practice

Problems 1 Draw

electron dot structures

for each molecule

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Covalent Bonding **Bonding Answer Key**

a bond formed when two atoms share a pair of electrons. Covalent bond. a chemical bond that involves sharing a pair of electrons betw.... Covalent compound. a chemical compound formed by the sharing of electrons. 14 terms. Tejas_Kode. Chemistry Chapter 16 Covalent Bonding. structural formulas.

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observe about the three structures? 2.
The following ...

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covalent. 2 atoms in
single covalent bond
share h.... 2 atoms in
double covalent bond
share h.... 2 atoms in
triple covalent bond
share h.... nonmetals
bond with each other

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by sharing electrons to
become....2.
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Bonding 5 Chapter 15 -
16 Assignment &
Problem Set 22. What
is an "alloy"? Chapter
16 Problems Single
Covalent Bonds 23.
Describe the difference
between an ionic and a
covalent bond. 24.
Classify the following

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- compounds as ionic or covalent: a. MgCl_2 b. Na_2S c. H_2O d. H_2S
25. Explain why neon is monatomic but chlorine is diatomic.
- 26.

Bonding Chapter 15 16 Assignment & Problem Set

Bonding Basics

Covalent Bonds

Answers This is

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Covalent Solids

Covalent solids A solid that consists of two- or three-dimensional networks of atoms held together by covalent bonds. are formed by networks or chains of atoms or molecules held together by covalent bonds. A perfect single crystal of

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a covalent solid is therefore a single giant molecule.

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to illustrate the

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Covalent
Bonding In covalent
bonds, electron sharing
usually occurs so that
atoms attain the

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electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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Single Covalent Bonds

When only one pair of electrons is shared,

such as in a hydrogen molecule, it is a single covalent bond. Chapter

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and group 16

elements. Using 1 and 16 to represent atoms of groups 1 and 16, respectively, the generic structure is:

Section 8.1

Assessment page 247

7. Identify the type of atom that generally forms covalent bonds.

The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule

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applies to covalent bonds.

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A bond formed by the sharing of electrons between atoms. A neutral group of atoms joined together by covalent bonds A molecule consisting of

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two atoms. The chemical bond that results from sharing valence electrons.... A chemical reaction in which energy is absorbed.

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